

# Kinetic and Mechanistic study of oxidation of isonicotinic acid hydrazide by Thallium (III) in acidic medium

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**Abstract** :The reaction between Thallium (III) and isonicotinic acid hydrazide is carried out in a mixture of perchloric and hydrochloric acid medium. The reaction proceeds through formation of complex with reactant, which decomposes in subsequent steps to give product. Effect of acrylonitrile shows, that there is no formation of free radicals. The increase in  $[H^+]$  and  $[Cl^-]$  decreases the rate of the reaction. The increase in ionic strength does not affect the rate of reaction. The effect of temperature was studied at four different temperatures ranging from  $15^{\circ}C$  to  $30^{\circ}C$ . The activation parameters were also determined and a mechanism is predicted.

**Keywords:** isonicotinic acid hydrazide, kinetics, thallium(III), oxidation

## Introduction

A detail investigation of oxidation of carboxylic acid hydrazide by thallium(III) have not been received much attention. Therefore, with a view to develop a new method, the present study was carried out. The reaction of hydrazides with most oxidants give the corresponding acids<sup>1</sup> and in some cases<sup>2</sup> esters or amides. Thallium(III) salts are well known oxidants<sup>3</sup> in organic synthesis. The oxidation of cyclopropane by thallium(III) acetate in acetic acid leads to cleavage of C-H bonds<sup>4</sup>.Thallium(III) can be utilized both as a strong (in perchloric acid and sulfuric acid media) and as a mild oxidant (in hydrochloric acid medium) by changing the reactive species. The chemical literature has described thallium-induced splitting of carbon-nitrogen bonds<sup>5</sup> no mechanistic investigation has been carried out. The present work deals with kinetic and mechanistic study of oxidation of isonicotinic acid hydrazide in hydrochloric acid medium.

## Experimental

Thallium (III) solution was prepared by dissolving  $Tl_2O_3$  (ACROS) in  $1.0 \text{ mol dm}^{-3}$  HCl and the concentration was ascertained by iodometric titration. The isonicotinic acid hydrazide was prepared from reported<sup>6</sup> procedure and characterized by determining its melting point. Stock solution of isonicotinic acid hydrazide was prepared in 50 % v/v, 1,4-dioxane. Ionic strength was kept constant.

The reactions were carried out in 50 % v/v 1-4 dioxane (s.d.fine.chem) under pseudo first order conditions keeping concentration of hydrazide in large excess over that of the oxidant. The solutions containing the reactants and all other constituents were thermally equilibrated separately, mixed and the reaction mixture was analysed for unreacted thallium (III) iodometrically by titrating against standard thiosulphate. The pseudo-first order rate constants were determined from the slopes of linear  $\log[Tl(III)]$  versus time plots. The results were reproducible up to  $\pm 5 \%$ . Kinetic runs were followed to about three half-lives of the reactions. Under the experimental condition oxidation of 1,4-dioxane did not occur.

The stoichiometry of the reaction was determined using a known excess of thallium (III) over hydrazide and determining remaining oxidant iodometrically after 24 hrs. The results consistent with equation-(1) were obtained. The corresponding carboxylic acid was characterized by determining its MP ( $310^{\circ}C$  for isonicotinic acid)



## Results

The reaction was carried out in a mixture of both the acids HCl and  $HClO_4$  because, the reaction occurs rapidly in perchloric acid medium but in the presence of hydrochloric acid the rate is measurable. The effect of reactants on the reaction was studied at constant  $[HCl]$  and  $[HClO_4]$  of  $0.1 \text{ mol dm}^{-3}$  each and ionic strength of  $0.6 \text{ mol dm}^{-3}$ .

### Effect of Hydrazide concentration

The effect of [hydrazide] was studied between  $1 \times 10^{-2}$  to  $1 \times 10^{-1}$   $\text{mol dm}^{-3}$ , keeping the [oxidant] constant at  $3.0 \times 10^{-3}$   $\text{mol dm}^{-3}$ . It is observed that the pseudo-first order rate constant increases, with increase in [hydrazide].

An examination of above result indicates that, the value of rate constant  $k$  depends on initial concentration of hydrazide and it increases with increase in hydrazide concentration.

### Effect of Thallium(III) concentration

To study the effect of Thallium (III) concentration on reaction rate, concentration of oxidant was varied from  $6.5 \times 10^{-4}$  to  $6.5 \times 10^{-3}$   $\text{mol dm}^{-3}$  keeping the [hydrazide] constant at  $1 \times 10^{-1}$   $\text{mole dm}^{-3}$ .

A persual of data shows that the rate of oxidation of isonicotinic acid hydrazide remains practically constant at five different concentrations of thallium (III). Hence, the order with respect to [oxidant] is unity.

### Effect of $\text{H}^+$ ion concentration

To study the effect of  $\text{H}^+$  ion concentration on the rate of oxidation of isonicotinic acid hydrazide, the concentration of HCl was varied and also different concentrations of NaCl were added in order to keep ionic strength constant.

A persual of the above data shows that, the rate of oxidation of isonicotinic acid decreases as the concentration of  $\text{H}^+$  ion increases.

### Effect of $\text{Cl}^-$ ion concentration

To study the effect of  $\text{Cl}^-$  ion concentration on the rate of oxidation of isonicotinic acid hydrazide by thallium (III), the concentration of HCl and  $\text{HClO}_4$  were varied keeping the concentration of other reactant constant.

From the above data, it shows that the rate of oxidation of isonicotinic acid hydrazide decreases as  $\text{Cl}^-$  ion concentration decreases.

### Effect of Temperature

In order to determine the temperature coefficient and thermodynamic parameters, the reaction was studied at four different temperatures ranging from  $15^\circ\text{C}$  to  $30^\circ\text{C}$  and at five different initial concentrations of hydrazide. A persual of the data clearly shows that the rates of oxidation of isonicotinic acid hydrazide are approximately doubled for  $10^\circ\text{C}$  rise in temperature.

### Effect of $\text{NaClO}_4$ concentration

In order to study the effect of  $\text{NaClO}_4$  on the oxidation of isonicotinic acid hydrazide, the concentration of  $\text{NaClO}_4$  solution was varied keeping the other reactants constant.

A persual of the data from table shows that the oxidation of isonicotinic acid hydrazide remains practically constant at five different concentrations of sodium perchlorate.

### Effect of Dielectric constant

The effect of dielectric constant ( $D$ ) of the medium on the rate of oxidation of isonicotinic acid hydrazide was studied by the addition of different volumes of ethanol (5% to 40%) to the reaction mixture, keeping the concentration of other reactants constant.

A persual of the data shows that as dielectric constant of the reaction mixture decreases rate of oxidation of isonicotinic acid hydrazide also decreases.

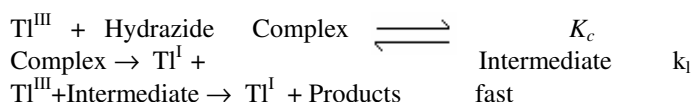
### Effect of Acrylonitrile

In order to determine whether free radical formation takes place during the course of oxidation it was necessary to determine the effect of acrylonitrile concentration on the rate of oxidation, the reaction was carried out at five different initial concentrations of acrylonitrile.

A persual of the data from table shows that the rate of oxidation of isonicotinic acid hydrazide remains practically constant at five different concentrations of acrylonitrile. This shows that there is no formation of free radicals in the reaction.

### Discussion

Since there is no formation of free radicals in the reaction, the reaction proceeds with two-electron transfer step. The order in thallium (III) was found to be unity and the order in hydrazide was found to be fractional.



### Scheme 1

Such fractional order in substrate concentration is due to the prior complex formation equilibrium between the reactants. The Michealis - Menten plots of  $1/k_{\text{obs}}$  versus  $1/[\text{Hydrazide}]$  were linear with an intercept in support of the complex formation. Therefore, in agreement with the results obtained the mechanism of the reaction can be represented as in Scheme 1. Equation 2 gives the rate according to Scheme1. Since, total  $[\text{TI}^{\text{III}}]$  exists in the form of free  $[\text{TI}^{\text{III}}]$  and the complex (Equation 3) therefore, the  $[\text{TI}^{\text{III}}]$  free is given by Equation 6. The overall rate law is now expressed by Equation 7 and the Pseudo-first order rate constant  $k_{\text{obs}}$ , by Equation 8.

$$\text{Rate} = k_1 [\text{Complex}] = k_1 K_c [\text{Hydrazide}]_{\text{free}} [\text{TI}^{\text{III}}]_{\text{free}} \quad (2)$$

$$[\text{TI}^{\text{III}}]_{\text{total}} = [\text{TI}^{\text{III}}]_{\text{free}} + [\text{Complex}] \quad (3)$$

$$[\text{TI}^{\text{III}}]_{\text{total}} = [\text{TI}^{\text{III}}]_{\text{free}} + K_c [\text{Hydrazide}] [\text{TI}^{\text{III}}]_{\text{free}} \quad (5)$$

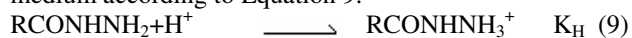
$$[\text{TI}^{\text{III}}]_{\text{free}} = [\text{TI}^{\text{III}}]_{\text{total}} / (1 + K_c [\text{Hydrazide}]) \quad (6)$$

$$\text{Rate} = k_1 K_c [\text{Hydrazide}] [\text{TI}^{\text{III}}]_{\text{free}} \quad (7)$$

$$k_{\text{obs}} = k_1 K_c [\text{Hydrazide}] / (1 + K_c [\text{Hydrazide}]) \quad (8)$$

Rate law 8 is verified by plotting  $1/k_{\text{obs}}$  against  $1/[\text{Hydrazide}]$  at four different temperatures and from the slopes and intercepts of these plots the values of  $k_1$  and  $K_c$  were calculated and are given in Table 1.

The effect of hydrogen and chloride ion concentrations on the reaction is due to the protonation of hydrazides<sup>7</sup> and different chloro – complexes of thallium (III) present in the solution. Hydrazides are known to be protonated in acid medium according to Equation 9.



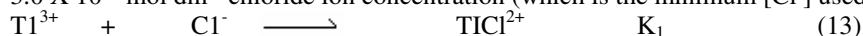
Therefore, total  $[\text{Hydrazide}]$  can be expressed by Equation 10 and thereby the fact that there was no effect of Free  $[\text{Hydrazide}]$  by Eq. 12. Since the rates of reaction decreases as the  $[\text{H}^+]$  increases, free hydrazide is the active species.

$$[\text{Hydrazide}]_{\text{total}} = [\text{Hydrazide}]_{\text{free}} + [\text{Hydrazide}]_{\text{protonated}} \quad (10)$$

$$[\text{Hydrazide}]_{\text{total}} = [\text{Hydrazide}]_{\text{free}} + K_H [\text{Hydrazide}]_{\text{free}} \quad (11)$$

$$[\text{Hydrazide}]_{\text{free}} = [\text{Hydrazide}]_{\text{total}} / (1 + K_H [\text{H}^+]) \quad (12)$$

Thallium (III) forms strong complexes with chloride ions of the formula  $\text{TlCl}_n^{3-n}$  where n is the number of chlorides complexes with thallium (III) as represented in equilibrium 13 to 16. The values of respective stability constants<sup>8</sup> are  $K_1 = 1.38 \times 10^8$ ,  $K_2 = 3.98 \times 10^{13}$ ,  $K_3 = 6.02 \times 10^{15}$  and  $K_4 = 1.0 \times 10^{18} \text{ mol}^{-1} \text{ dm}^3$ . The presence of  $3.0 \times 10^{-2} \text{ mol dm}^{-3}$  chloride ion concentration (which is the minimum  $[\text{Cl}^-]$  used in the present study).



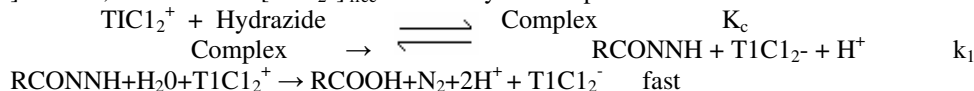
All the thallium (III) will exists as  $\text{TlCl}_2^+$  and its concentration can be expressed by Equation 17. The  $[\text{TlCl}_2^+]_{\text{free}}$  can now be given by eq. 19 where,  $\beta_1 = K_3/K_2 = 151$  and  $\beta_2 = K_4/K_3 = 166$ , further, using Equations 18 and 19 the concentrations of  $[\text{TlCl}_2^+]_{\text{free}}$ ,  $\text{TlCl}_3$  and  $\text{TlCl}_4^-$  were calculated at different chloride ion concentrations and compared with the change in rate constant as the chloride ion concentration varied.

$$[\text{TI (III)}]_{\text{total}} = [\text{TlCl}_2^+]_{\text{total}} = [\text{TlCl}_2^+]_{\text{free}} + [\text{TlCl}_3] + [\text{TlCl}_4] \quad (17)$$

$$[\text{TlCl}_2^+]_{\text{total}} = [\text{TlCl}_2^+]_{\text{free}} (1 + \beta_1 [\text{Cl}^-] + \beta_2 [\text{Cl}^-]^2) \quad (18)$$

$$[\text{TlCl}_2^+]_{\text{free}} = [\text{TlCl}_2^+]_{\text{total}} / (1 + \beta_1 [\text{Cl}^-] + \beta_2 [\text{Cl}^-]^2) \quad (19)$$

The concentration of both of  $[\text{TlCl}_2^+]_{\text{free}}$  and  $\text{TlCl}_3$  parallel the values of rate constants as  $[\text{Cl}^-]$  changes but the order  $[\text{Cl}^-]$  is – 1.5, which makes  $[\text{TlCl}_2^+]_{\text{free}}$  as the only active species.



where  $\text{R} = \text{C}_5\text{H}_4$  for isonicotinic acid hydrazide

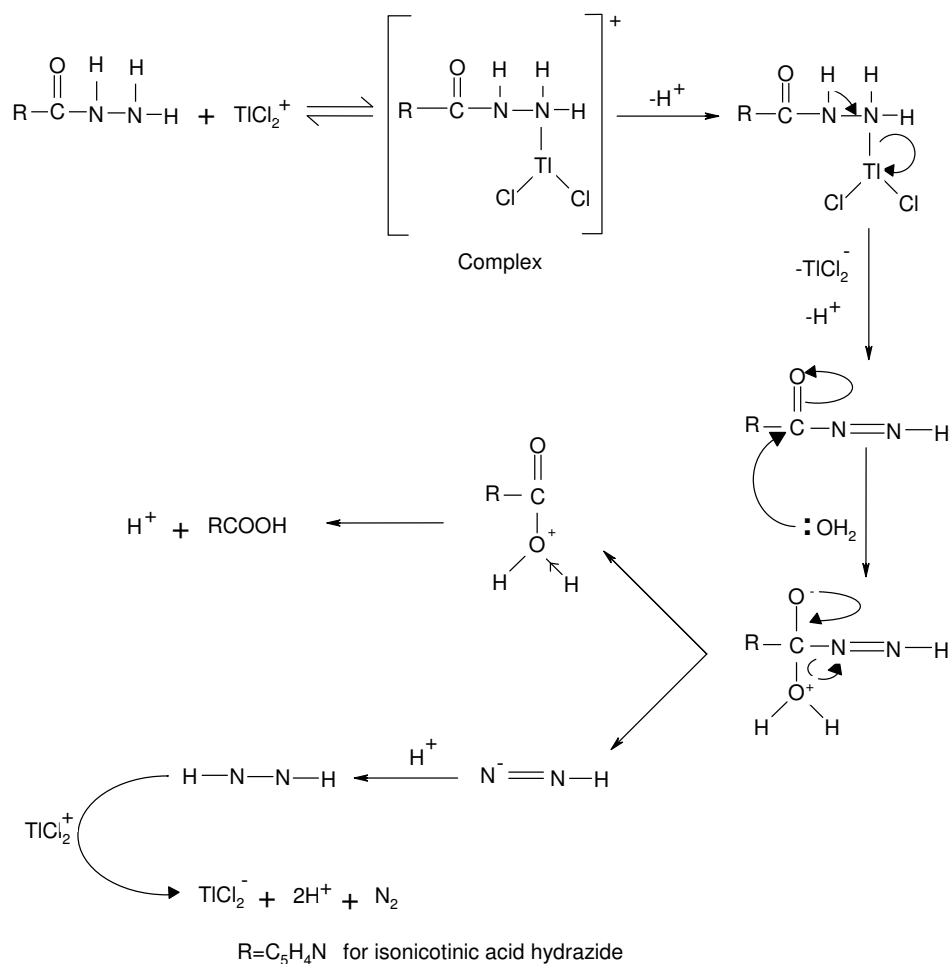
### Scheme 2

The mechanism considering  $\text{TlCl}_2^+$  of oxidant and free hydrazide of the substrate as the active species can now be represented by scheme 2 with respective rate law and the expression for the pseudo-first order rate constants by Equations 20 and 21. The rate law 21 was verified by plotting  $1/k_{\text{obs}}$  against  $1/[\text{Hydrazide}]$  and  $1/k_{\text{obs}}$  against  $[\text{H}^+]$  which were found to be linear. From the slopes and intercepts of these plots the values of  $K_c$  and  $K_H$  were determined.

$$\text{Rate} = \frac{k_1 K_c [\text{Hydrazide}]_{\text{total}} [\text{TlCl}_2^+]_{\text{total}}}{(1 + K_c [\text{Hydrazide}]) (1 + K_H [\text{H}^+]) (1 + \alpha_1 [\text{Cl}^-] + \alpha_2 [\text{Cl}^-]^2)} \quad (20)$$

$$k_{\text{obs}} = \frac{k_1 K_c [\text{Hydrazide}]_{\text{total}}}{(1 + K_c [\text{Hydrazide}]) (1 + K_H [\text{H}^+]) (1 + \alpha_1 [\text{Cl}^-] + \alpha_2 [\text{Cl}^-]^2)} \quad (21)$$

The values of  $K_c$  are given in Table 1 and those of  $K_H$  were found to be 13 and 16  $\text{mol}^{-1} \text{dm}^3$  for isonicotinic acid. The electrophilic character of  $\text{TlCl}_2^+$  among the thallium (III) chlorocomplexes is highest thus making it the reactive species.



**Scheme 3**

The detailed mechanism involves electrophilic substitution on the nitrogen of the hydrazide with the formation of N-Tl bond, which decomposes in the subsequent step with, direct two-electron transfer from hydrazide to thallium to give an intermediate followed by fast steps. (Scheme 3). Such N-Tl bond formation has been postulated during thallium (III) oxidation of nitrogen<sup>9</sup> containing compounds.

The activation parameters, with respect to slow step,  $k_1$ ,  $\Delta H^\ddagger$  ( $\text{KJ mol}^{-1}$ ),  $\Delta G^\ddagger$  ( $\text{KJ mol}^{-1}$ ) and  $\Delta S^\ddagger$  ( $\text{JK}^{-1} \text{mol}^{-1}$ ) were found to be 59.74, 90.42 and -102.96 respectively for isonicotinic acid hydrazide. Considerable decrease in the entropy of activation is due to formation of more ordered transition state as shown in scheme 3. The mechanism involves neutral hydrazide as the active substrate thus the reaction is unaffected by the change in the ionic strength. The increase in ethanol content in the reaction medium decreases; the rate such an effect of the solvent is due to the stabilization of the complex formed between reactants<sup>10</sup> in a medium of low relative permittivity.

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**Table 1 – values of K<sub>c</sub> and k<sub>1</sub>**

[HCl] = 0.1 mol dm<sup>-3</sup>, [HClO<sub>4</sub>] = 0.1 mol dm<sup>-3</sup>

[Tl<sup>III</sup>] = 3.0 X 10<sup>-3</sup> mol dm<sup>-3</sup>,

I = 0.6 mol dm<sup>-3</sup> in 50% v/v 1,4-Dioxane

Temperature °C	K <sub>c</sub> (mol <sup>-1</sup> dm <sup>3</sup> ) [Isoniazid]	k <sub>1</sub> X 10 <sup>-4</sup> (S <sup>-1</sup> ) [Isoniazid]
15	6.17	10.0
20	7.52	14.3
25	8.1	20.0
30	9.7	33.34

## References

- 1) a) Ho T.L, Ho H. C. and Wong C. M., Synthesis., 1972, 562.  
b) Schnyder J. and Rittenberg M., Helv. Chim. Acta., 1975, 58, 521.  
c) Tsuji J., Takayanagi H and Toshida Y., Chem. Lett., 1976, 147.
- 2) a) Aylward J. B., Norman R. O. C, J. Chem. Soc.(C)., 1968, 2399.  
b) Tsuji J. Hayakawa H. Takayanagi, Chem. Lett., 1975, 437.  
c) Tsuji J., Nagashima T., Qui, N. T and Takayanagi H, Tetrahedron., 1980, 36, 1311  
d) Hoffman R. V., Kumar A. J., J. Org. Chem., 1984, 49, 4014.
- 3) McKillop A. Taylor E. C, Organic synthesis by oxidation with metal compounds., W. J. Mijs, C. R. H. J. de Jonge, Plenum: Newyork, 1989, 5<sup>th</sup> edn., p 695.
- 4) South Jr. A. and Ouellette R. J, J. Am. Chem. Soc., 1980, 90, 7064
- 5) (a) McKillop A. Hunt J. D and Taylor E. C, J. Am. Chem. Soc., 1971, 93, 4918.  
b) Butler, R. N. Morris G. J and O'Donohue A. M, J. Chem. Res.(S)., 1981, 61.  
c) Taylor E. C., Robey R. L., McKillop A. A., Angew.Chem. Int. Ed., Engl., 1972, 11, 48.  
d) Taylor E. C, Robey R. L., McKillop A. A, J. Org. Chem., 1972, 37, 2797.  
e). Silveria Jr. A, Angelastro M., Israel, R. Totino F and Williamsen, P. J. J. Org. Chem., 1980, 45, 3522.
- 6) . Vogel A. I, Textbook of practical organic chemistry., (ELBS & Longman Group) 4<sup>th</sup> Edn 1975, p 1125.
- 7) (a) Kazo K., Hirakazo T., Hisashi K and Zenzo T., Chem. Pharm. Bull., 1963, 11, 797.  
(b) Krishnarao P. V., Frank M. S. and Ramaih, A. K. *React. Kinet. Cata. Lett.*, 1978, 9, 59.  
(c) Krishnarao P. V, Frank M. S. and Ramaih A.K Indian J. Chem., 1978, 16 A, 418.  
(d) Ramaih A. K, Frank M. S, Baburao G. and Krishnarao. P. V. Indian J. Chem., 1979, 18 A, 416.
- 8) Lee A. G. The Chemistry of Thallium., Elsevier C., London, 1971, p 48.
- 9) (a) McKillop J. D Hunt R. D. and Taylor E. C., J. Am. Chem. Soc., 1971, 93, 4918.  
(b) Butler R. N, Morris G. J. and O' Donohue A. M., J. Chem. Res. (s)., 1981, 61.
- 10) Amis E. S, Solvent effects on reaction rates and mechanisms (Academic Pres., New York) 1966.

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